

INVESTIGATION OF FACTORS AFFECTING THMs FORMATION IN DRINKING WATER

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ABSTRACT

Total Trihalomethanes (THMs) content can vary greatly, depending both on water quality and on treatment conditions, in the same water supply system. Most municipal water treatment plants use chlorine to keep water biologically safe during the distribution process. The effect of various parameters on the formation of THMs has been studied in this investigation using water samples collected from Zai water treatment plant in Jordan. These parameters are reaction time, temperature, pH, TOC, free chlorine, and bromide ion content. An increase in the value of any of these parameters has positive effects on the formation of THMs. A general correlation for the prediction of THMs as a function of these parameters is proposed. Its predictions are in good agreement with the observed results regardless of the fact that it is restricted to this study.

Keywords: THMs formation, chlorination, drinking water, water quality model.

INTRODUCTION

Chlorine, while an effective disinfectant, reacts with organic materials that might be present in the source drinking water to produce a group of chlorinated organic compounds known as Disinfection Byproducts (DBPs). The DBPs include THMs (chloroform, bromodichloromethane, dibromochloromethane, and bromoform). The presence of THMs in drinking water has been the subject of increasing public concern since the 1970s. Toxicological studies suggested that THMs have been linked to the occurrence of human cancer in many instances (US National Cancer Institute, 1976). This was soon followed by an act by the Environmental Protection Agency (EPA) to regulate their concentrations in drinking water.

In the past, various investigations had been conducted to determine factors that might influence THMs formation (Richardson, 2005). From the mid of 1980s, several models simulating the propagation and transformation of contaminants in a distribution system over time were developed and applied (Grayman *et al.*, 1988), (Liou and Kroon, 1987). In spite of water treatment, finished water may sustain substantial changes in quality while being transmitted through the distribution system before reaching the consumer's tap. These changes may be associated with complex physical, chemical and biological processes that take place during transport.

Garcia-Villanova *et al.* (1997) developed a statistical correlation that predicts the concentration of chloroform formation as function of pH, temperature and distance (near or far) between a reservoir and a point of interest in the distribution system. This correlation is qualitative in the sense that distance and time are not included implicitly in it.

Golfinopoulos *et al.* (1998) generated a multiple regression model for predicting THMs level in the finished water leaving a treatment plant. The correlation implies that THMs formation is higher in summer through autumn and lower in winter through spring.

Clark (1998) described a mathematical model that applies second order kinetics to predict the chlorine residuals and demonstrates that THMs formation can be characterized as a function of chlorine demand. The parameters of this model are correlated as a function of pH, TOC, temperature and chlorine residual.

The formation of THMs in drinking water has been shown to be a function of various water quality parameters and chlorination conditions. THMs formation is one of the slowest known reactions in

water treatment and a number of important factors influence this reaction. These include residence time, pH, temperature, bromide ion concentration, chlorine dosage, type of organic precursor and the Total Organic Carbon (TOC).

In the present paper, the overall objective is to develop an improved understanding of the effect of water parameters on THMs formation in drinking water, and also develop a mathematical model for the prediction of THMs concentrations for Zai water treatment plant.

MATERIALS AND METHOD

The experiments in this study were conducted on samples taken from Zai water treatment plant located near the capital Amman in Jordan. The procedure is based on collecting 2 liters samples from the feed and the effluent streams of the plant.

The samples were stored at a given constant temperature. For THMs measurement smaller samples were taken from the original sample in 40 ml septum vials containing few crystals of sodium thiosulfate (sufficient to eliminate any residual chlorine and thus stop the trihalomethane formation reaction), filled to zero head space, and stored in the dark at the desired experimental conditions prior to analysis. The THMs were measured using Head Space Analyzer GC with ECD.

The experimental conditions, other than the parameter to be investigated, were held constant at some predetermined values. The following parameters were investigated.

The reaction time was held between 2-3 hrs. This is usually sufficient for the reaction to go to completion within one hour which is typical for these reactions.

Effect of Reaction Time

The concentrations of THMs were determined after the following reaction times have elapsed: 0, 1.25, 5.25, 10.0, 18.0, 23.0, 29.25, and 49.25 hr. The other parameters were fixed at: temperature 16°C, pH 7.8, TOC 1.14 ppm, free chlorine 1.04 ppm, and bromide ion 0.42 ppb. Trihalomethanes formation potential was determined after 168 hr (seven days).

Effect of Temperature

The effect of temperature on the formation of THMs was carried out at temperatures of 5, 14, 28, and 40 °C. The other parameters were fixed at: TOC 0.829 ppm, pH 7.71, free chlorine 1.2 ppm, reaction time 3 hr, and bromide ion 0.5 ppb.

Effect of pH

The pH was checked and changed for each sample by the drop-wise addition of 0.1 M HCl or 0.1 M NaOH. The pH values used in this part of the investigation were: 6.04, 7.15, 7.71 and 8.7. The pH was measured using an ORION 960 Automatic System, with pH electrode (model No. 9172 BN).

The other experimental conditions were fixed at temperature 20°C, TOC 0.829 ppm, free chlorine 1.15 ppm, reaction time 2.5 hr, and bromide ion 0.5 ppb.

Effect of Chlorine Dose

The residual chlorine checked and changed for each sample by the drop wise addition of standard chlorine solution prepared from 6% sodium hypochlorite solution. The DPD-colorimetric method was employed for residual chlorine measurement

The concentrations of THMs were determined at the following conditions: temperature 14°C, TOC 0.829 ppm, pH 7.71, reaction time 1.5 hr, and bromide ion 0.5 ppb, and at the following chlorine residuals values (1.2, 1.75 and 2.88 ppm).

Effect of TOC

In order to study the effect of TOC different values in the lab, drops of influent water were added to the effluent water to achieve the required TOC concentration for each final sample. The following TOC values: 1.33, 1.82, and 2.04 mg/l, were considered in this study. The TOC analyzer employed was a DC-180 (3300) (Dorhman / Rosemount) with Non Dispersive Infrared Detector (NDIR).

With these TOC levels, the concentrations of THMs were determined for water samples under the following conditions: reaction time 2 hrs, temperature 15°C and bromide ion concentration of 0.55 ppb. In these experiments there were very slight variations in pH and chlorine residuals for the three samples. These changes in pH are due to the presence of organic acids resulting in pH values of: 7.84, 8.01, 7.88. Rapid chlorine decay reaction resulted in different levels of chlorine: for chlorine: 0.84, 0.7, 0.6 ppm. For those reasons an average pH value of 7.91 and an average chlorine residual of 0.74 mg/l were taken for those samples.

Effect of Bromide Ion

The effect of the presence of bromide ions on the formation of THMs was demonstrated by measuring the bromide concentration in water samples using an ion chromatographic method. A DX-120 Ion Chromatograph (DIONEX) with chemical suppressor for anions and conductivity detector was used for this purpose.

RESULTS AND DISCUSSION

General Correlation

In order to describe the mechanisms of THM formation, several empirical models have been proposed (Golfinopoulos *et al*, 1998), (Garcia-Vollanova *et al*, 1997), (El-Dib, 1995). The models developed range from over all reaction stoichiometry expressions to empirical equations derived from linear and nonlinear statistical regression analysis. These equations predict the concentration of THMs as function of the various parameters under consideration.

In this study the following predictive equation is proposed:

$$THMs = 105.4958t^{0.2705} (6.72 - 1.733pH + 0.135 pH^2) T^{(0.00665 T)} 0.59^{Cl_2} Cl_2^{1.3588} TOC^{(0.99/70C)} Br^{1.69} \quad (1)$$

in which:

THMs: total trihalomethnes concentration, ppb

t: time, hours

pH: value of pH in solution

T: temperature, °C

Cl₂: chlorine concentration, ppm

Br₂: bromine concentration, ppm

TOC: total organic carbon level, ppm

The parameters of this equation were determined by applying standard regression techniques to the experimental data obtained in this study. The goodness of fit of the data predicted by this correlation to experimental data is demonstrated in Figure (1). The coefficient of correlation being 0.982. However, this predictive equation is only applicable to this study or any other study with the same conditions. This was concluded after many trials to apply this predictive equation to another data from different country.

Effect of Residence Time

The formation of THMs under natural conditions is not instantaneous. Under some reaction conditions the formation of THMs may be completed in less than an hour, in other circumstances, several days may be required before the maximum yield of THMs occurs.

Concentration levels of THMs formed after chlorination of drinking water issuing from Zai water treatment plant at various residence times are presented in Figure (2). It is seen that THMs concentrations increase as the residence time

extends from 1.25 hr to 49.25 hr. The increase is initially rapid slowing down as time goes on. This trend resembles a first order behavior. The Figure also illustrates the good agreement between experimental and predicted data.

Analytical kinetic models of THMs formation have not yet been developed, mainly because of the complexity of the reactions of chlorine with organic materials in water. The consumption of chlorine in water occurs through complex sequential and parallel reaction pathways with unknown rates (Kavanaugh, 1980).

It was proposed in the EPA model that THMs formation follows first order growth limited reaction, for which the trihalomethane formation potential (CFP) is the limiting concentration to which THMs can grow.

This model is given by:

$$\frac{dC}{dt} = k(C_{FP} - C) \quad (2)$$

which upon integration yields the following equation:

$$C = C_o + (C_{FP} - C)[1 - e^{(-kt)}] \quad (3)$$

where:

C: total trihalomethnes concentration, ppb

t: time, hours

C_o: initial THMs concentration, ppb

C_{FP}: THMs formation potential, ppb

K: rate constant, hr⁻¹

This equation can be rearranged into the following form:

$$\ln\left(\frac{C_{FP} - C_o}{C_{FP} - C}\right) = kt \quad (4)$$

This is a straight line equation passing through the origin and its slope is the rate constant k. A plot of the left hand side of this equation based on experimental data versus time is shown in Figure (3). It is very clear that the data obtained in this study follow first order kinetics since the plot gives a straight line. The rate constant k is found using standard regression methods to obtain the slope of this line. The rate obtained is 0.0022 hr⁻¹.

Effect of Temperature

Data collected by the Water Authority of Jordan (Garber, 1991) showed that THMs concentration varied seasonally and were lower in the winter and higher in summer due to the effect of temperature variation from winter to summer.

With chlorination of natural water, approximately twice as much of chloroform can be formed during a given

period of time at 25°C as is formed at 3°C (Rook, 1979).

Increasing the temperature has a positive effect on THM formation as shown in Figure (4). It is also seen that this trend is very well predicted by the correlation proposed in this study.

Effect of pH

The reaction sequence of THMs formation suggests that the pH would have a strong effect on THMs formation. This effect has been illustrated by a number of researchers (Kavanaugh, 1980), (Rook, 1979), (El-Dib, 1995).

Rook (1979) demonstrated that a large increase in the reaction rate occurred at pH values from 8-10; however, this effect was weakly pronounced in the range 6-8. He illustrated the increase at the high pH values was the result of increased formation of phenoxide ions. The results by El-Dib (1995) also showed that the reaction rate of THMs progressively increased as the pH increased from 6 to 9.

In the present study, it is observed that THMs concentrations increase with increasing pH as shown in Figure (5). There again the good agreement between predictions and experimental data is shown on the same Figure.

Effect of Chlorine Dose

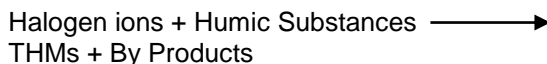
Trussell and Umphres (1978) described the chlorine reaction with THMs precursors as a three-step reaction. The first step was where immediate chlorine demand was exerted. Typically this is due to the reaction with an inorganic rather than an organic compound. Once the immediate demand is satisfied, the additional chlorine begins to react with any available organic material, including THM precursors. They showed that in this step there was a nearly linear relationship between chlorine dose and the concentration of THMs produced. When enough chlorine was consumed satisfying both the immediate and short-term organic chlorine demands, a long-term chlorine residual is obtained, and any further increase in THM formation with chlorine dosage is small.

Kavanaugh (1980) showed that the rate of THMs formation is strongly dependent on the applied chlorine dose, i.e. it increases with increasing chlorine dose.

As shown in Figure (6), it is noticed that, as expected, the THMs concentration increased as a function of the applied free chlorine dose. Also, the Figure shows the predicted results as compared to the experimental data.

Effect of Total Organic Carbon

Trihalomethanes are formed according to the following generalized equation (Edzwald, 1987):



The structure of humic substance is proposed by Vasconcelos (1996); while the reaction mechanism was proposed by Morris and Baum (1997).

The formation of THM is strongly dependent on the amount of naturally present humic substances (Humic and Fulvic acid), generally referred to as Total Organic Carbon (TOC). i.e. it increases with increasing TOC in presence of free chlorine residuals.

As shown in Figure (7), it is seen that THMs concentration increases almost linearly with TOC concentration. This trend is very well followed by the predictive equation as shown in the Figure.

Effect of bromide ion

Both chlorine and bromine are oxidants and substitution agents. Chlorine being a better oxidizing agent; while bromine being a better substitution agent. Moreover, the kinetics of bromine substitution is more rapid than those for chlorine substitution. This is why bromide ions were found to be good precursors in the formation of brominated THMs in the presence of chlorine (even with the presence of substantial amounts of it). Bromide ions are oxidized to bromine (hypobromous acid) which participate in the halogenation step in the THMs formation reaction sequence much more effectively than does chlorine.

The results of the measurements made to reveal the effect of the presence of bromide ions on the THMs formation in the Zai water treatment plant are shown in Figure (8). They clearly indicate that the presence of bromide ions has significantly enhanced the formation of brominated THM species in agreement with results found in the literature (Toroz et al, 2005). The formation of Bromoform is by large dominating amongst the halogenated methanes formed.

CONCLUSIONS

The results of this investigation indicated that increasing the residence time, temperature, pH, free chlorine, and TOC (humic substances) had a positive effect on THM formation; i.e. THMs concentration increased as the values of these parameters increased. These results were found to be compatible with literature findings. The results also showed that the presence of bromide ions in the chlorinated drinking water increased the formation of brominated-THM species, especially bromoform.

It is shown in this study that the factors affecting the formation of THMs can be related by simple regression correlation based on experimental data. The correlation proposed in this study satisfactorily predicted the THMs formation as a function of the parameters under investigation.

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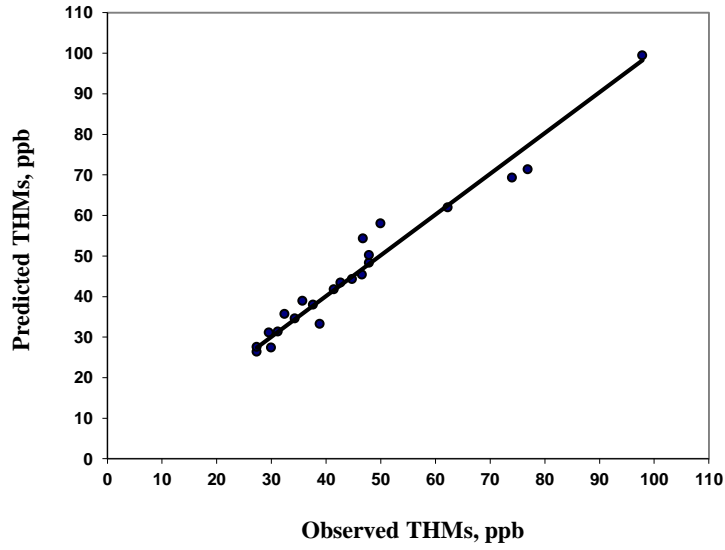


Figure (1): Comparison between Observed and Correlated Data of THMs

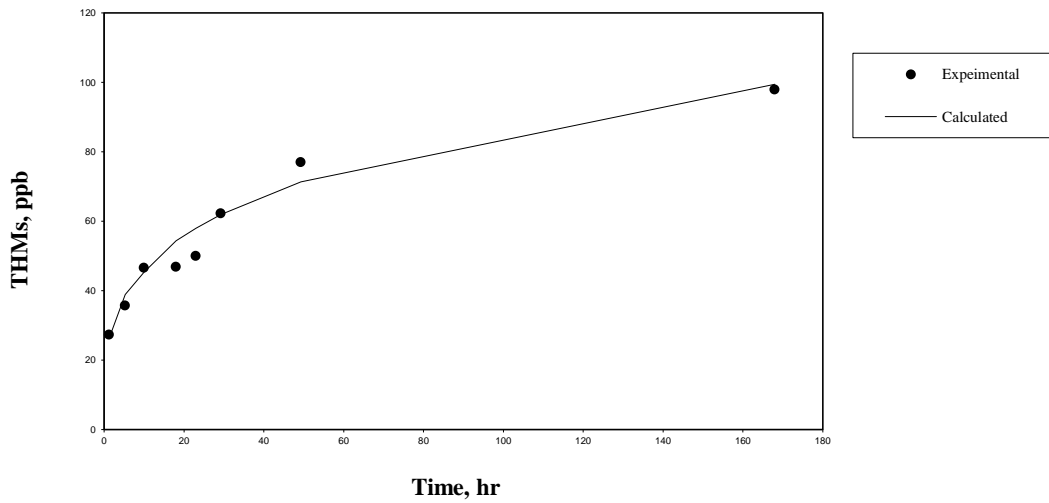
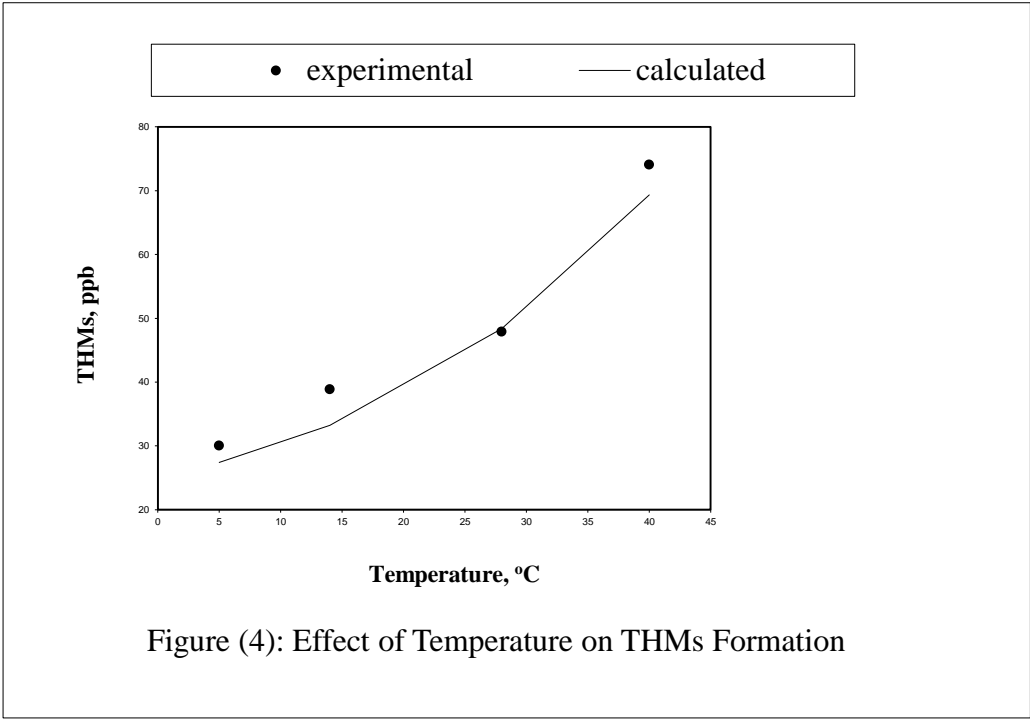
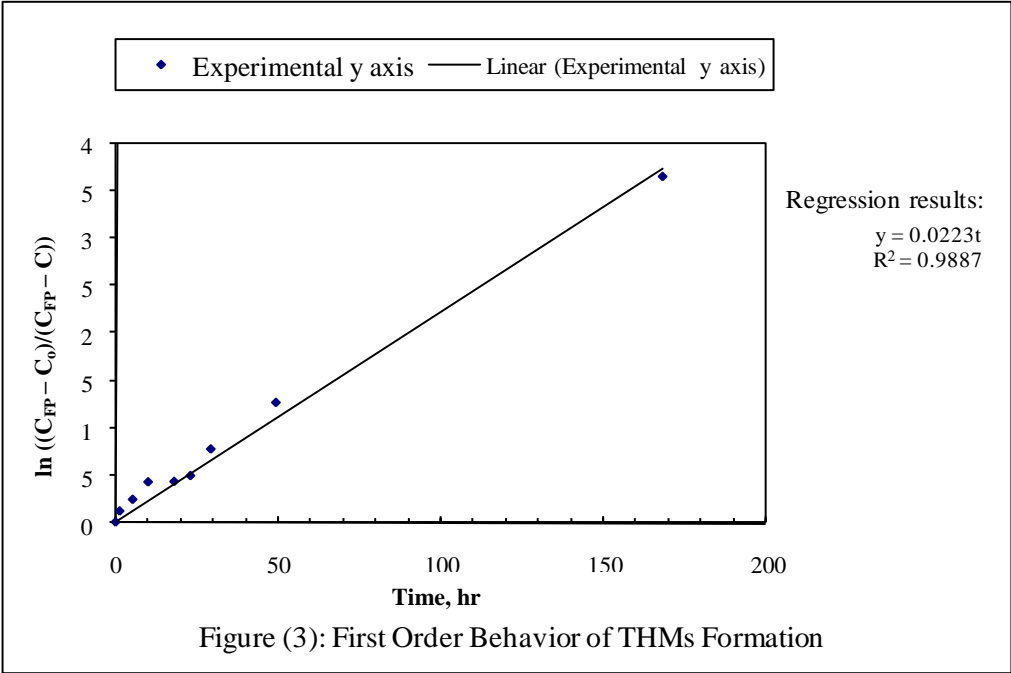


Figure (2): Effect of Reaction Time on THMs Formation



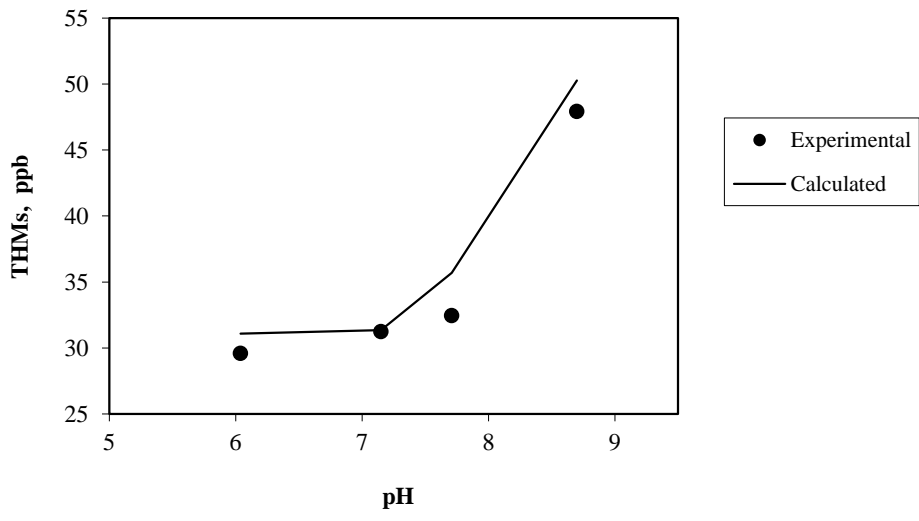


Figure (5): Effect of pH on THMs Formation

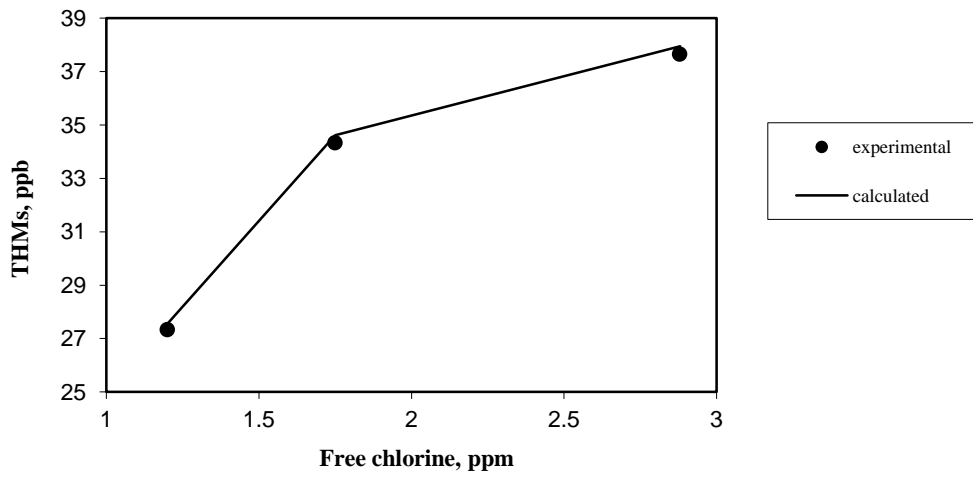


Figure (6): Effect of Chlorine on THMs Formation

